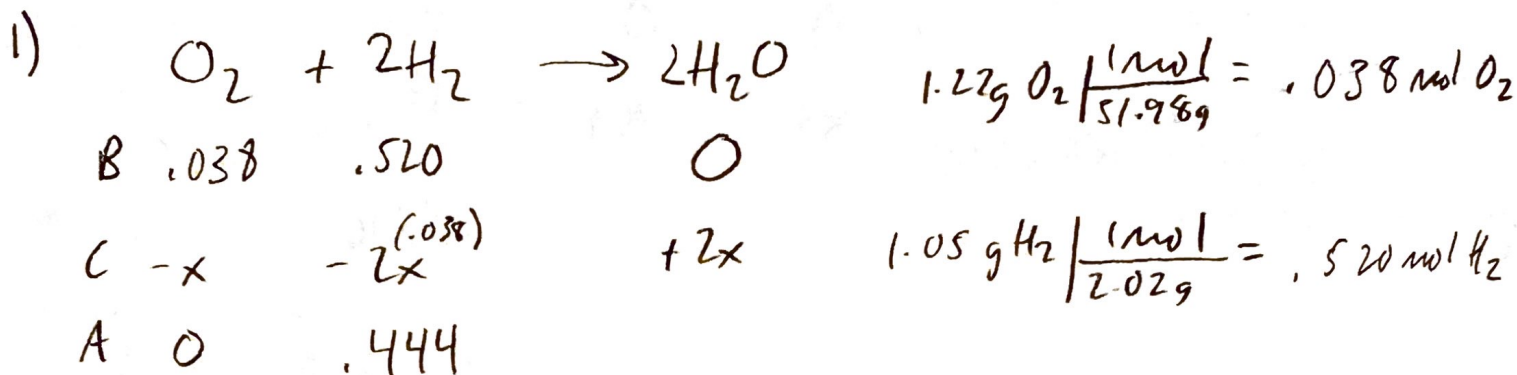
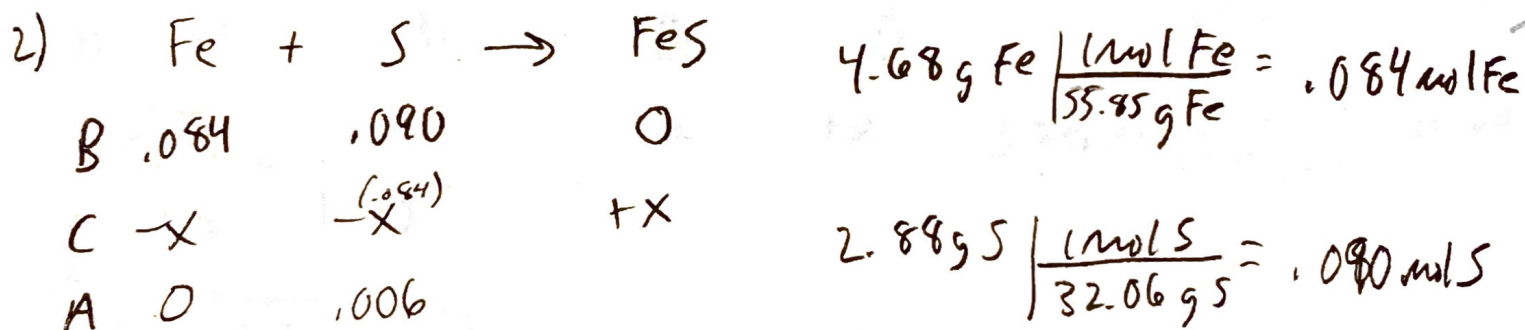


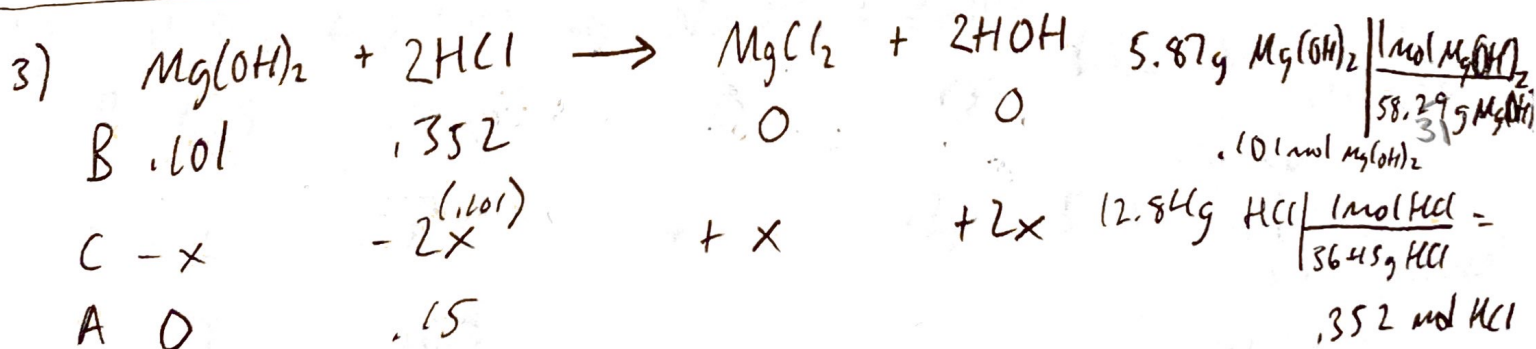
1-3 Practice Problems



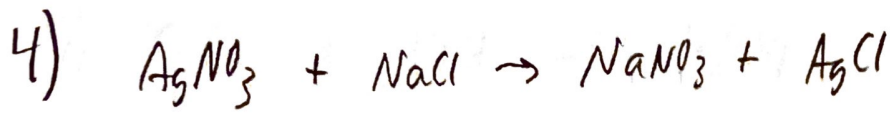
LR = O₂



LR = Fe



LR = Mg(OH)₂



$$B \quad .051 \quad .070 \quad 0 \quad 0$$

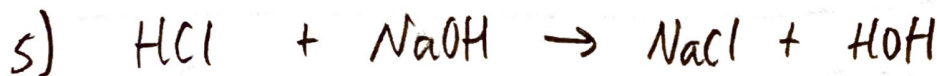
$$C \quad -x \quad -x^{(.051)} \quad +x \quad +x$$

$$A \quad 0 \quad .019$$

$$6.25 \text{ g AgNO}_3 \left| \frac{1 \text{ mol AgNO}_3}{171.86 \text{ g AgNO}_3} = .036 \text{ mol AgNO}_3$$

$$4.12 \text{ g NaCl} \left| \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} = .070 \text{ mol NaCl}$$

$$\boxed{LR = \text{AgNO}_3}$$



$$B \quad .214 \quad .131 \quad 0 \quad 0$$

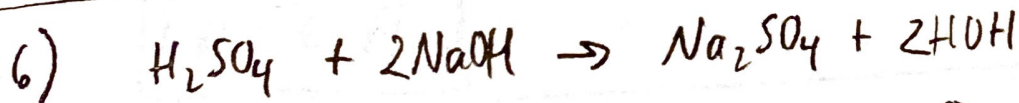
$$C \quad -x^{(.131)} \quad -x \quad +x \quad +x$$

$$A \quad .083 \quad 0$$

$$7.81 \text{ g HCl} \left| \frac{1 \text{ mol HCl}}{36.45 \text{ g HCl}} = .214 \text{ mol}$$

$$5.24 \text{ g NaOH} \left| \frac{1 \text{ mol NaOH}}{39.98 \text{ g NaOH}} = .131 \text{ mol}$$

$$\boxed{LR = \text{NaOH}}$$



$$B \quad .065 \quad .148 \quad 0 \quad 0$$

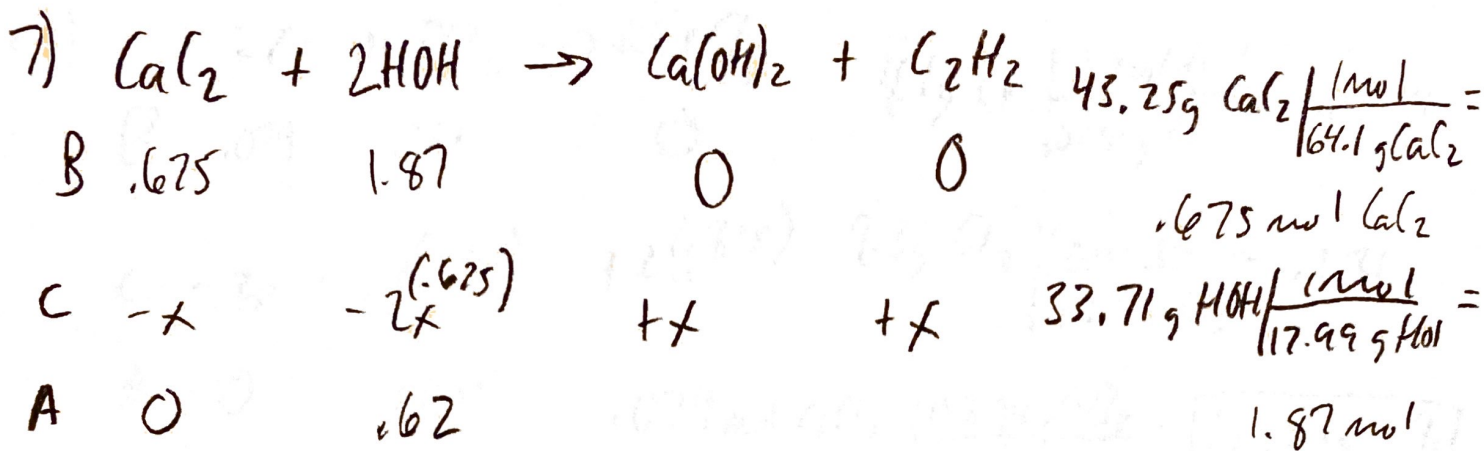
$$C \quad -x \quad -2x^{(.065)} \quad +x \quad +2x$$

$$A \quad 0 \quad .018$$

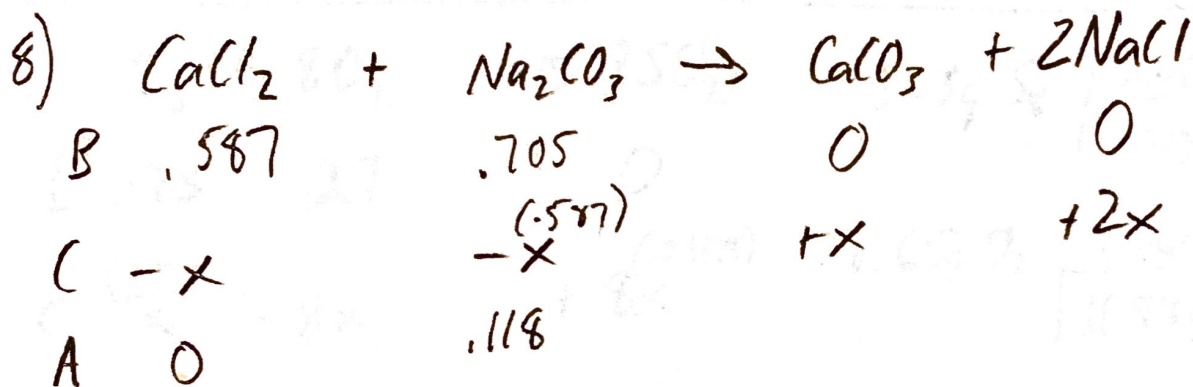
$$6.33 \text{ g H}_2\text{SO}_4 \left| \frac{1 \text{ mol}}{98.02 \text{ g H}_2\text{SO}_4} = .065 \text{ mol}$$

$$5.92 \text{ g NaOH} \left| \frac{1 \text{ mol}}{39.98 \text{ g NaOH}} = .148 \text{ mol}$$

$$\boxed{LR = \text{H}_2\text{SO}_4}$$



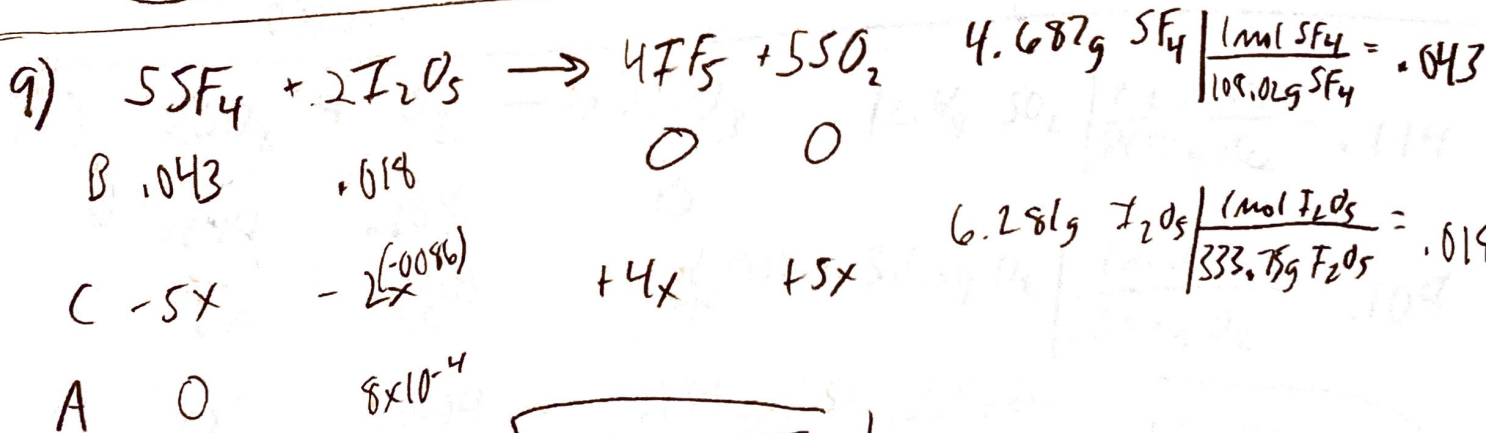
LR = CaI₂



$65.14\text{g CaCl}_2 \left| \frac{1\text{mol CaCl}_2}{110.98\text{g CaCl}_2} = .587$

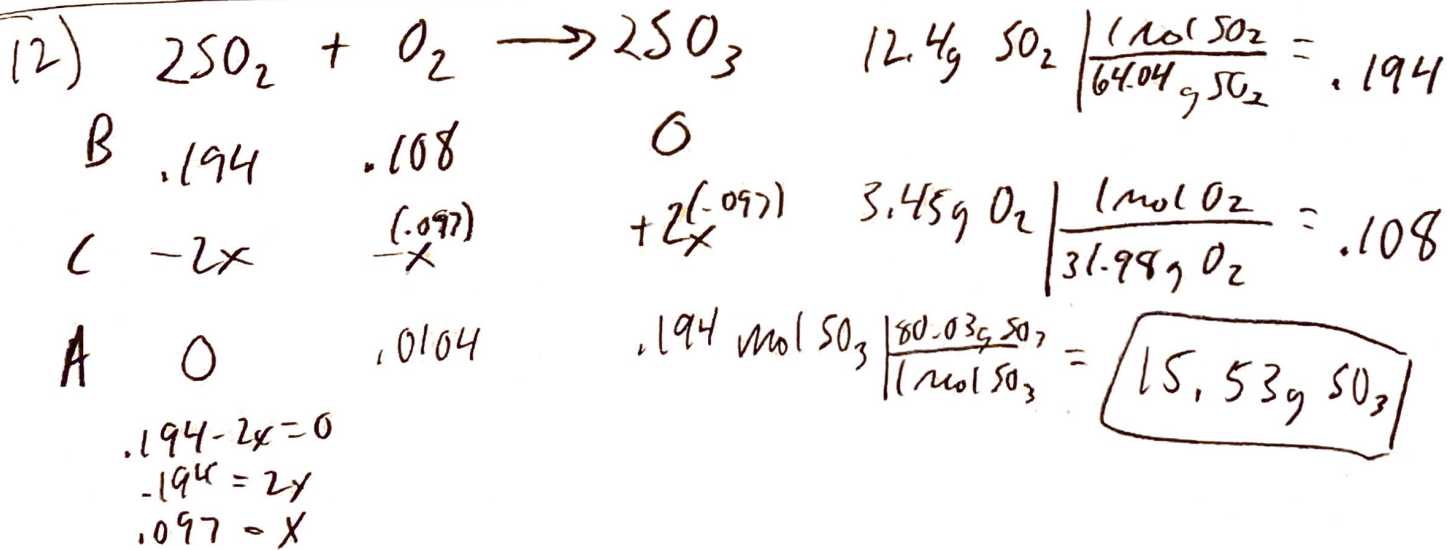
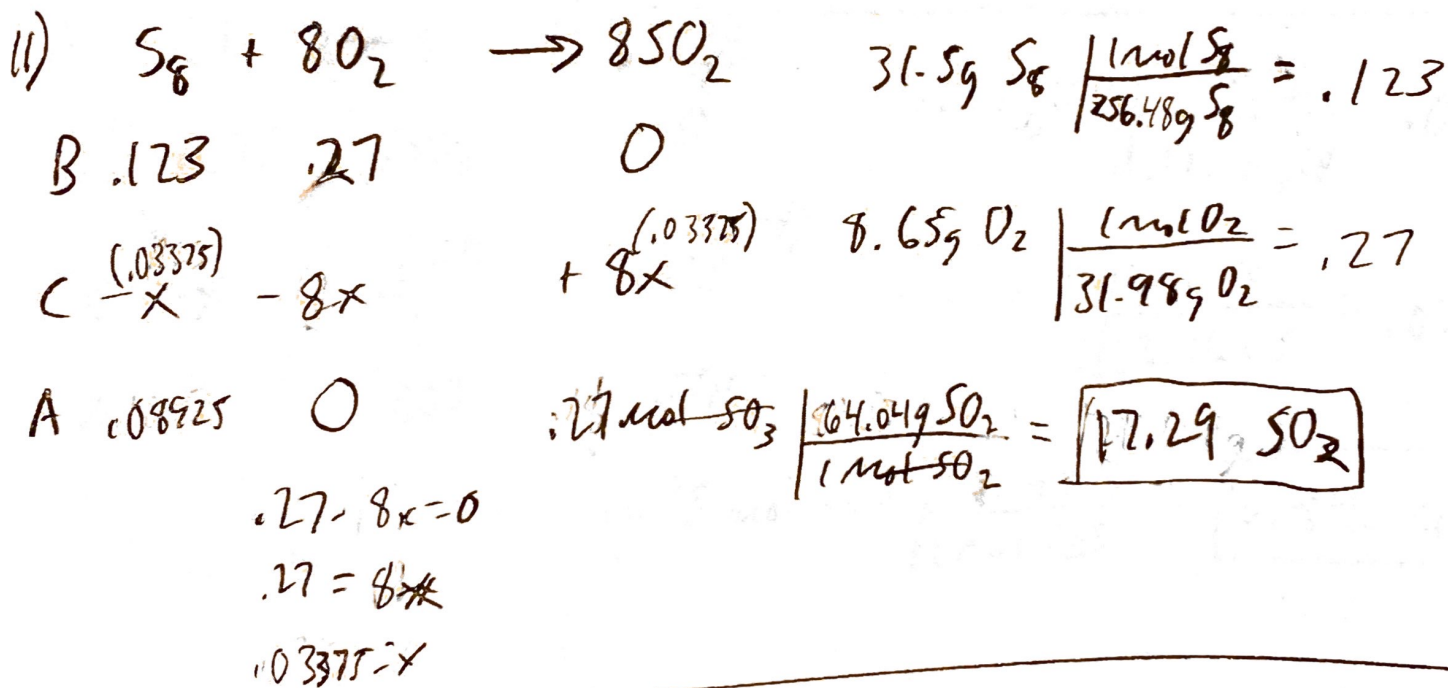
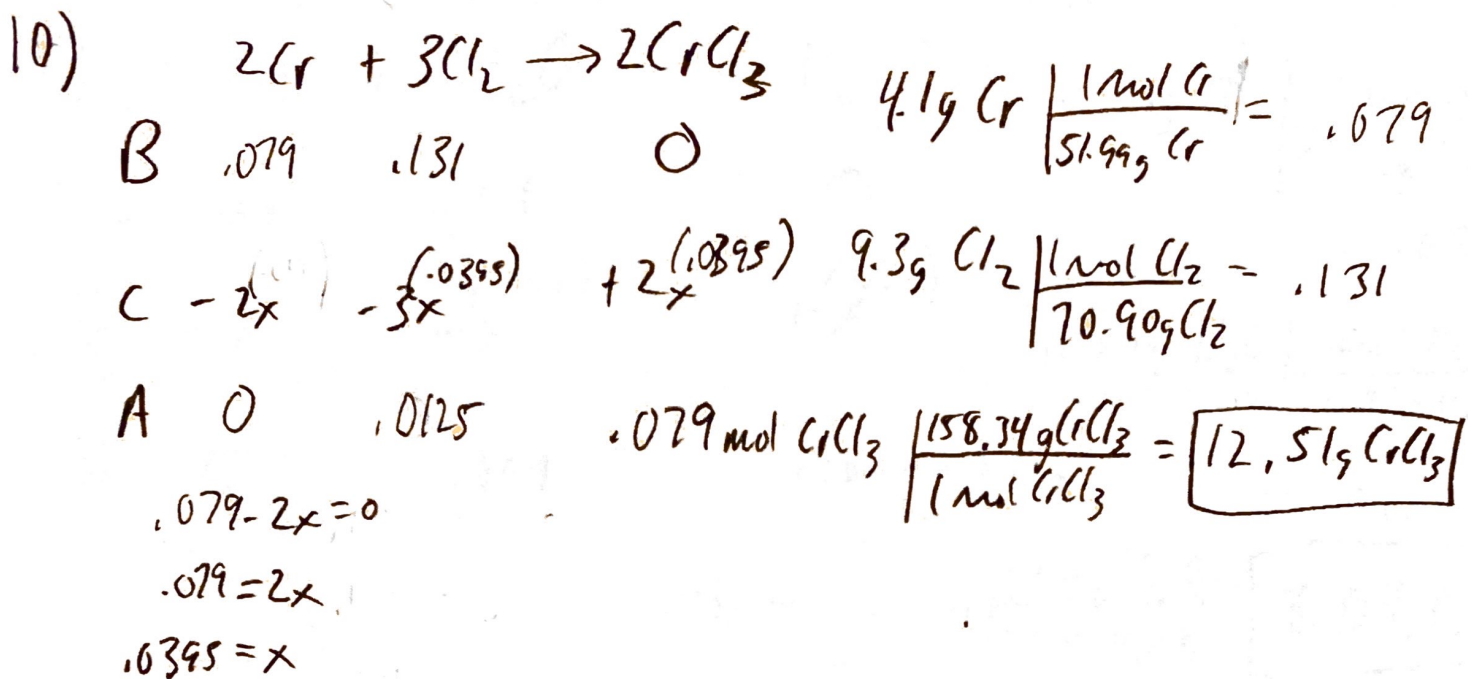
$74.68\text{g Na}_2\text{CO}_3 \left| \frac{1\text{mol Na}_2\text{CO}_3}{105.96\text{g Na}_2\text{CO}_3} = .705$

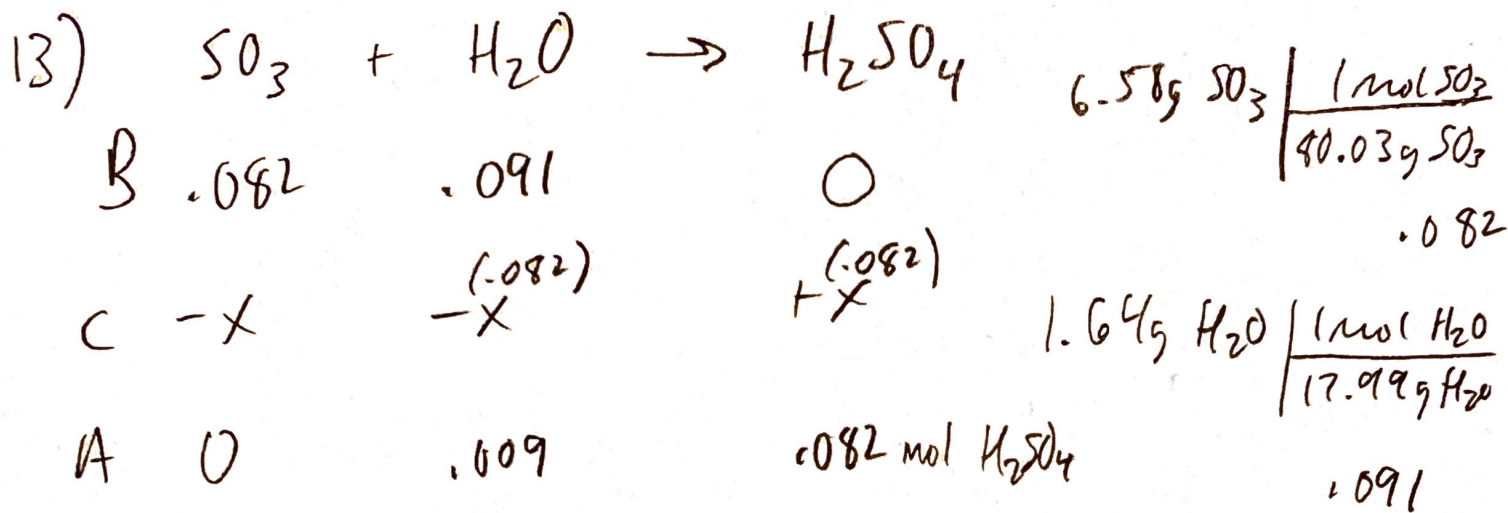
LR = CaCl₂



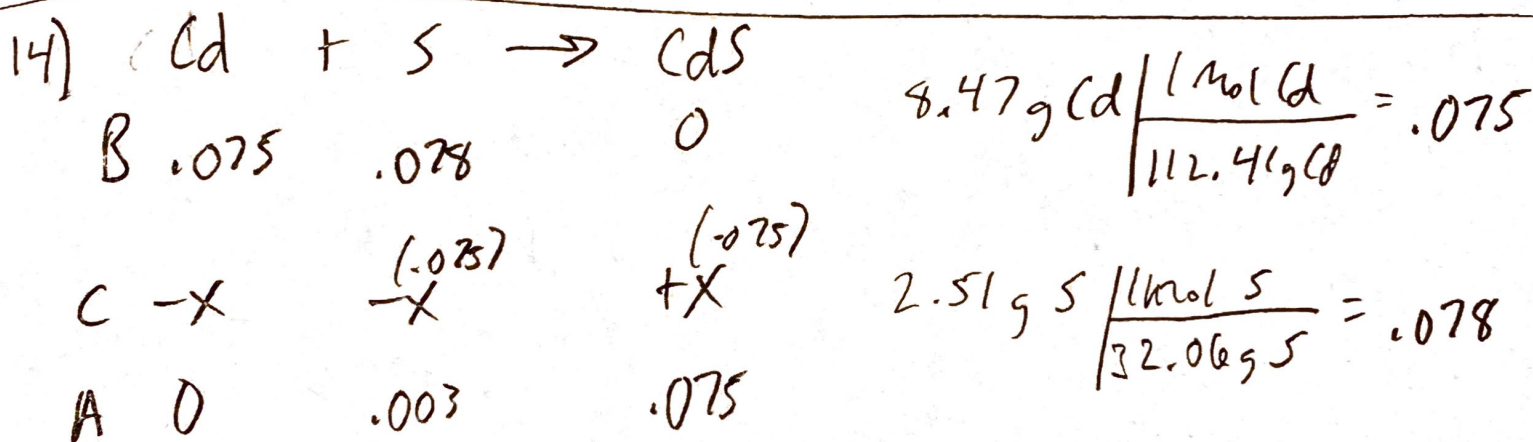
LR = SF₄

$.043 - 5x = 0$
 $.043 = 5x$
 $.0086 = x$





$$.082 \text{ mol H}_2\text{SO}_4 \left| \frac{98.02\text{g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \right. = \boxed{8.04\text{g H}_2\text{SO}_4}$$



$$.075 \text{ mol CdS} \left| \frac{144.47\text{g CdS}}{1 \text{ mol CdS}} \right. = \boxed{10.84\text{g CdS}}$$