

Ions in Chemical Compounds

Name _____

Complete the following table, being sure that the total charge on the resulting compound is zero.

<u>Ions</u>	Chloride Cl	Hydroxide OH	Nitrate NO ₃	Sulfide S	Carbonate CO ₃	Phosphate PO ₄
Sodium Na						
Ammonium NH ₄						
Potassium K						
Calcium Ca						
Magnesium Mg						
Aluminum Al						
Iron (II) Fe						
Iron (III) Fe						
Copper (I) Cu						

Naming Ionic Compounds

Give the name and molar mass of the following ionic compounds:

	Name	Molar Mass
1)	Na_2CO_3	_____
2)	NaOH	_____
3)	MgBr_2	_____
4)	KCl	_____
5)	FeCl_2	_____
6)	FeCl_3	_____
7)	$\text{Zn}(\text{OH})_2$	_____
8)	Be_2SO_4	_____
9)	CrF_2	_____
10)	Al_2S_3	_____
11)	PbO	_____
12)	Li_3PO_4	_____
13)	TiI_4	_____
14)	Co_3N_2	_____
15)	Mg_3P_2	_____
16)	$\text{Ga}(\text{NO}_2)_3$	_____
17)	Ag_2SO_3	_____
18)	NH_4OH	_____
19)	$\text{Al}(\text{CN})_3$	_____
20)	$\text{Be}(\text{CH}_3\text{COO})_2$	_____

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For the following compounds, give the formulas and the molar masses:

	Formula	Molar Mass
22)	sodium phosphide	_____
23)	magnesium nitrate	_____
24)	lead (II) sulfite	_____
25)	calcium phosphate	_____
26)	ammonium sulfate	_____
27)	silver cyanide	_____
28)	aluminum sulfide	_____
29)	beryllium chloride	_____
30)	copper (I) arsenide	_____
31)	iron (III) oxide	_____
32)	gallium nitride	_____
33)	iron (II) bromide	_____
34)	vanadium (V) phosphate	_____
35)	calcium oxide	_____
36)	magnesium acetate	_____
37)	aluminum sulfate	_____
38)	copper (I) carbonate	_____
39)	barium oxide	_____
40)	ammonium sulfite	_____
41)	silver bromide	_____
42)	lead (IV) nitrite	_____

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