

Name _____
 Period ____ Date ____/____/____

• Bonding & Molecular Structure

LEWIS STRUCTURES

Indicate the # of VALENCE electrons for each species. Write the correct Lewis electron-dot structure for each.

F # of valence e ⁻ 's = ____ 	O # of valence e ⁻ 's = ____ 	K # of valence e ⁻ 's = ____ 	Al # of valence e ⁻ 's = ____
F⁻ # of valence e ⁻ 's = ____ 	O²⁻ # of valence e ⁻ 's = ____ 	K⁺ # of valence e ⁻ 's = ____ 	Al³⁺ # of valence e ⁻ 's = ____
F₂ # of valence e ⁻ 's = ____ 	H₂ # of valence e ⁻ 's = ____ 	HF # of valence e ⁻ 's = ____ 	NH₃ # of valence e ⁻ 's = ____
CH₄ # of valence e ⁻ 's = ____ 	NF₃ # of valence e ⁻ 's = ____ 	SiF₄ # of valence e ⁻ 's = ____ 	C₂H₆ # of valence e ⁻ 's = ____
MgH₂ # of valence e ⁻ 's = ____ 	LiH # of valence e ⁻ 's = ____ 	AlH₃ # of valence e ⁻ 's = ____ 	BH₃ # of valence e ⁻ 's = ____

C₂H₄ # of valence e ⁻ 's = _____	C₂F₄ # of valence e ⁻ 's = _____	CO # of valence e ⁻ 's = _____	O₂ # of valence e ⁻ 's = _____
CO₂ # of valence e ⁻ 's = _____	C₂H₂ (H C C H) # of valence e ⁻ 's = _____	N₂ # of valence e ⁻ 's = _____	HCN # of valence e ⁻ 's = _____
CN⁻ # of valence e ⁻ 's = _____	SO₄²⁻ # of valence e ⁻ 's = _____	PO₄³⁻ # of valence e ⁻ 's = _____	ClO₃⁻ # of valence e ⁻ 's = _____
CO₃²⁻ # of valence e ⁻ 's = _____	NO₃⁻ # of valence e ⁻ 's = _____	SO₂ # of valence e ⁻ 's = _____	O₃ (O O O) # of valence e ⁻ 's = _____
SF₆ # of valence e ⁻ 's = _____	XeF₄ # of valence e ⁻ 's = _____	PCl₅ # of valence e ⁻ 's = _____	SeF₄ # of valence e ⁻ 's = _____