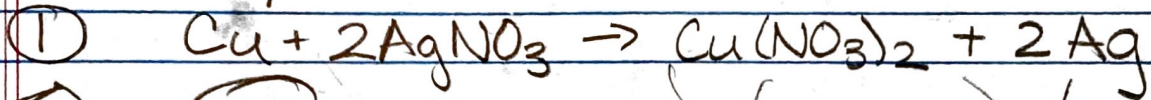
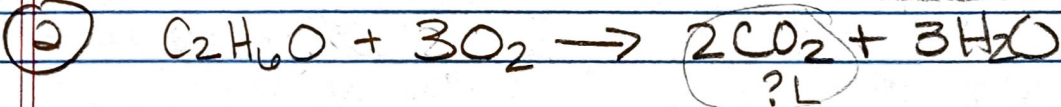


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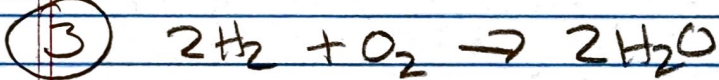


B	2.9	5.8	0	0
C	-X	-2X (2.9)	+X	+2X
A	0	0	0	0

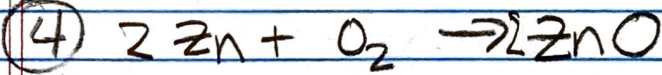


B	5.420	0	0	0
C	-X	-3X	+2X (5.420)	+3X
A	0	0	10.84 mol	22.4L CO ₂

1 mol = 242.82L CO₂



B	2500	0	0
C	-2X	-X	+2X (2500)
A	0	0	5 mol H ₂ O 18g H ₂ O = 90g H ₂ O



B	2.20	0	0
C	-2X	-X (2.20)	+2X = 4.41
A	0	0	2.20 mol 32g O ₂ = 70.4g O ₂

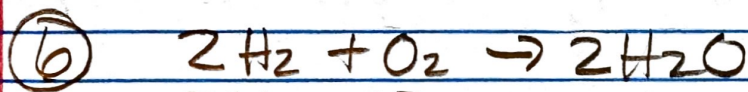
358.5g ZnO | 1 mol = 81.38g ZnO



B	5.56	?	8.34	0	0
C	-2X	-3X	+X	+3X	0
A	X=2.78	3(2.78)	0	0	0

150g Al | 1 mol Al = 5.56 mol | 26.98g Al

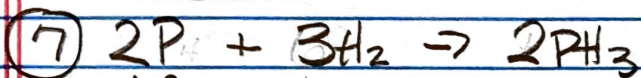
8.34 mol H₂SO₄ | 98.07g = 817.90g H₂SO₄



	3.6L	L?	
B	0.161	.081	0
C	-2X	-X(.081)	+2X
A			

$$\frac{3.6\text{L H}_2}{22.4\text{L}} \Big| \frac{1\text{mol}}{2} = 0.161$$

$$2X = \frac{.161}{2} = .081\text{mol} \Big| \frac{22.4\text{L O}_2}{1\text{mol}} = 1.8\text{L O}_2$$

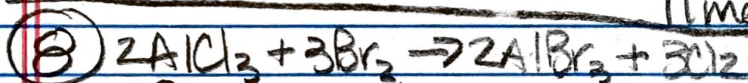


	6.2g	4.0g	
B	.200	2 mol	0
C	-2X(.1)	-3X(.1)	+2X(.1)
A	0	XS	?

$$\frac{6.2\text{g P}}{30.97\text{g P}} \Big| \frac{1\text{mol}}{2} = 0.200\text{mol P}$$

$$\frac{4\text{g H}_2}{2\text{g H}_2} \Big| \frac{1\text{mol}}{3} = 2\text{mol H}_2$$

$$.2\text{mol PH}_3 \Big| \frac{33.97\text{g PH}_3}{1\text{mol}} = 6.79\text{g PH}_3$$



	84.2g	68.4g		
B	.632	.428	0	0
C	-2X	-3X	+2X	+3X(.143)
A	XS	.413		.429

$$\frac{84.2\text{g AlCl}_3}{133.33\text{g AlCl}_3} \Big| \frac{1\text{mol}}{2} = .632$$

$$\frac{68.4\text{g Br}_2}{159.8\text{g Br}_2} \Big| \frac{1\text{mol}}{3} = .428$$

$$.429\text{mol Cl}_2 \Big| \frac{70.9\text{g Cl}_2}{1\text{mol Cl}_2} = 30.42\text{g Cl}_2$$

⑨