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| 1 | **Name: (NH4)2SO4** |
| 2 | **Name: N2O3** |
| 3 | **Name: PCl5** |
| 4 | **Name: Fe2O3** |
| 5 | **Name: SF6** |
| 6 | **Name: HNO3** |
| 7 | **Name: H3N** |
| 8 | **Name: LiCl** |
| 9 | **Name: FeS** |
| 10 | **Name: HNO2** |
| 11 | **Name: PbO** |
| 12 | **Name: N2O5** |
| 13 | **Name: H2SO3** |
| 14 | **Write the formula: calcium oxide** |
| 15 | **Write the formula: hydroiodic acid** |
| 16 | **Write the formula: dinitrogen tetraoxide** |
| 17 | **Write the formula: lead (IV) sulfate** |
| 18 | **Write the formula: dinitrogen monoxide** |
| 19 | **Write the formula: Copper (II) sulfate** |
| 20 | **How many total atoms are in 2Li2SO3?** |
| 21 | **Define ionic bond: (talk about charge, what the electrons are doing)** |
| 22 | **Define covalent bond: (talk about what the electrons are doing)** |
| 23 | **Which bond type melts at a higher temperature?** |
| 24 | **Which bond type conducts electricity?** |
| 25 | **Balance & identify type: N2 + H2 → NH3** |
| 26 | **Balance & identify type : Li + AlCl3 → LiCl + Al** |
| 27 | **Balance & identify type: N2O5 → N2 + O2** |
| 28 | **Identify type (Honors balance): C10H24 + O2 → CO2 + H2O** |
| 29 | **Balance & identify type: NH4OH + H3PO4 → (NH4)3PO4 + H2O** |
| 30 | **Ionic and covalent bonds change their electrons so they have a full \_\_\_\_?** |

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| 31 | **Complete: Copper (II) Chloride + Potassium Nitrate🡪** |
| 32 | **Complete: Tin + Aluminum Oxide →** |
| 33 | **Complete: Iron(III) Chloride + ammonium hydroxide 🡪** |
| 34 | **Complete: Potassium Hydroxide + Calcium Fluoride →** |
| 35 | **Complete: Cesium metal + Nitrogen gas →** |
| 36 | **Complete: C6H8 + O2 →** |
| 37 | **Complete: Zn + CuO →** |
| 38 | **Complete: Aluminum Chloride →** |
| 39 | **Draw Nitrate: Which is the geometry, polarity, bond angle, and IMF’s present?** |
| 40 | **List the IMF’s in increasing order. What is the trend for IMF’s and boiling point?** |

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| # | **Answer (this part will be graded)** |
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